

PEER REVIEW

**LEMBAR
HASIL PENILAIAN SEJAWAT SEBIDANG ATAU PEER REVIEW
KARYA ILMIAH : JURNAL ILMIAH**

Judul Jurnal Ilmiah (Artikel) : Quantitatives Structure-Property Relationship Analysis Againts Critical Micelle Concentration of Sulfonate-Based Surfactant Based on Semiempirical Zindo/1 Calculation

Penulis Jurnal Ilmiah* : Ponco Iswanto, Eva Vaulina YD, Ely Setiawan, Fiandy Aminullah Putra

Jumlah Penulis : 4 Orang

Status Penulis : Penulis ke-1

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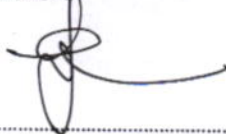
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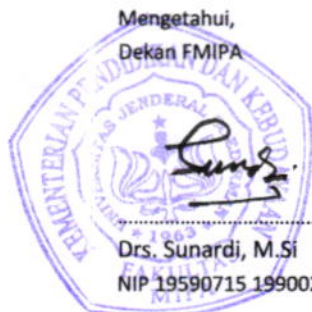
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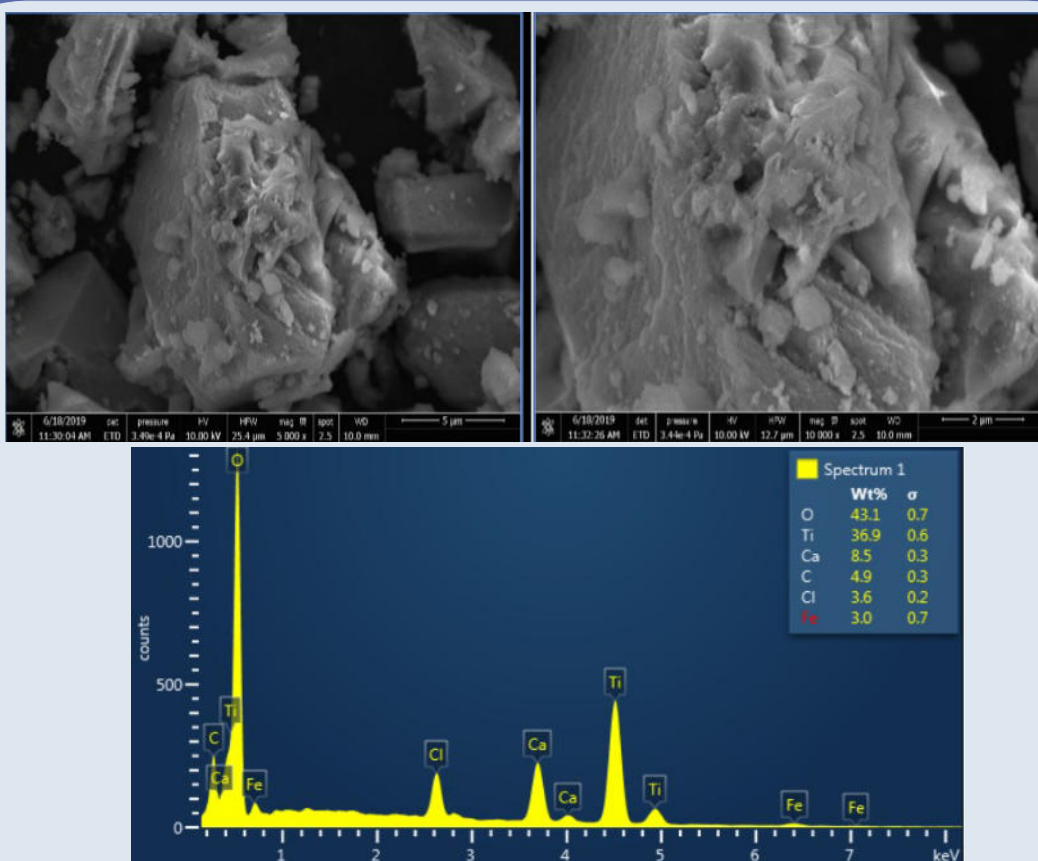
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ARTIKEL

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PREFACE

MOLEKUL, a scientific journal in chemistry, constantly publishes articles related to the chemistry of regional focus and are related to the environment or natural products of Indonesia. We published the journal in electronic version at the website of www.jmolekul.com, which was indexed by SCOPUS, CAS (Chemical Abstract Service), ASEAN Citation Index, DOAJ, CrossRef and Sinta.

In this volume 14 issue 2, natural products of Indonesia have been reported such as *Plectranthus Amboinicus* (Lour.) Spreng., *Tetracera indica*, *Apis dorsata*, *Archidendron jiringa* (Jack) I.C Nielsen, *Pimpinella pruatjan* Molk., *Aglalia harmsiana*, *Durio zibethinus* L., *Brassica juncea* and *Helianthus annuus* L..

We would like to thank all authors and reviewers for their contributions.

Editor-in-Chief
Amin Fatoni

TABLE OF CONTENT

Quantitative Structure-Property Relationship Analysis Against Critical Micelle Concentration of Sulfonate-Based Surfactant Based on Semiempirical Zindo/1 Calculation Ponco Iswanto, Eva Vaulina Yulistia Delsy, Ely Setiawan, Fiandy Aminullah Putra	78 - 83
Optimization of Total Flavonoids Extraction and α-Glucosidase Inhibitory Activity from <i>Plectranthus Amboinicus</i> (Lour.) Spreng. Leaves using the Simplex-Centroid Design Kunto Wibisono, Syarifah Iis Aisyah, Sri Suhesti, Waras Nurcholis	84 - 91
The bioactive contents and antioxidant activity of honey bee nest extract of <i>Apis dorsata</i> Binghami from the North Sulawesi Mokosuli Yermia Samuel, Eva Sherly Nonke Kaunang, Jacklin Stella Manopo	92 - 102
β-Sitosterol and Betulonic Acid from n-Hexane Extract the Stem Bark of <i>Tetracera indica</i> Muharni, Elfita, Heni Yohandini, Julinar, Muthia Oktaviani	103 - 109
Effectiveness of Leaves Extract Fractions of <i>Archidendron jiringa</i> (Jack) I.C Nielsen Against Microbes Oom Komala, Sri Wardatun, Lia Puspita Sari	110 - 116
No Acute Toxicity Tests of Purwoceng (<i>Pimpinella pruatjan</i> Molk.) Ethanolic Extract on Male Albino Rat by Determined Hepatorenal Function Test and Histopathology Fitranto Arjadi, Dhadhang Wahyu Kurniawan, Yudhi Wibowo, Wahyu Siswandari, Lantip Rujito	117 - 125
Caryophyllene-Type Sesquiterpenoids from the Stembark of <i>Aglalia harmsiana</i> and Their Cytotoxic Activity Against MCF-7 Breast Cancer Cells Hersa Milawati, Desi Harneti, Rani Maharani, Nurlelasari, Ace Tatang Hidayat, Mohamad Nurul Azmi, Yoshihito Shiono, Unang Supratman	126 - 132
Xylose and Arabinose Fermentation to Produce Ethanol by Isolated Yeasts from Durian (<i>Durio zibethinus</i> L.) Fruit Hermansyah, Fachrijal, Miksusanti, Fatma, Getari Kasmiarti, Almunadi T. Panagan	133 - 139
Synthesis and Characterization of TiO_2-CaO and TiO_2-CaO-Fe_2O_3 Photocatalyst for Removal of Catechol Candra Yulius Tahya, Wahyu Irawati, Karnelasatri, Friska Juliana Purba	140 - 148

Computational Approach to Evaluate Eugenol Affinity and Derivatives Empirical Against Cu(II)

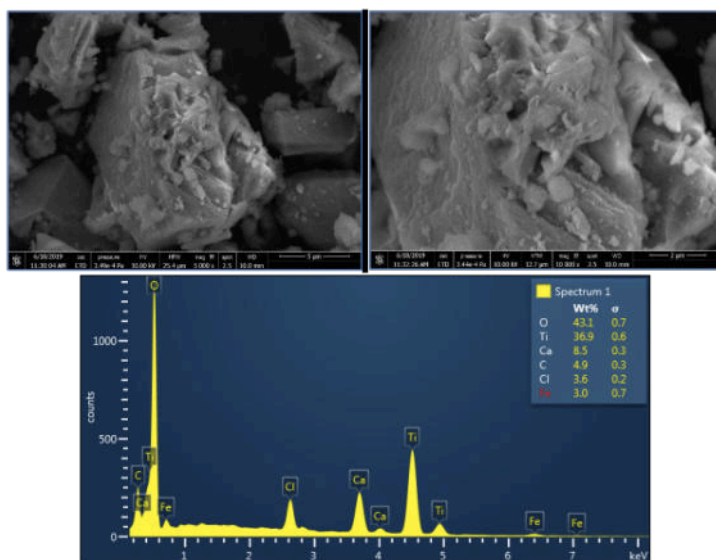
Muhammad Cholid Djunaidi, Tantri Nevi Astuti, Parsaoran Siahaan

149 - 156

Toxicity test of desorbed chromium from polluted soil around tannery industry in Yogyakarta Indonesia by using *Brassica juncea* and *Helianthus annuus* L.

Suherman, Wina Akmarina, Wulandari Sutiawan, Nurul Hidayat Aprilita, Mudasir, Kinichi Morita

157 - 163



Cover figure:

SEM and EDS results of TiO₂-CaO-Fe₂O₃ catalyst
(see Chandra Yulius T., et.al., pp 140-148)

Quantitative Structure-Property Relationship Analysis Against Critical Micelle Concentration of Sulfonate-Based Surfactant Based on Semiempirical ZINDO/1 CalculationPonco Iswanto^{1*}, Eva Vaulina Yulistia Delsy¹, Ely Setiawan¹,
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ABSTRACT. Development of anionic surfactant compound is very important because the anionic surfactant class is widely used in people's lives. For instance, anionic surfactants are used as food additives and detergents. The novel compound of sulfonate-based surfactant or proposed compound has predicted the Critical Micelle Concentration (CMC) value of experiment. Quantitative Structure-Property Relationship (QSPR) analysis based on semiempirical ZINDO/1 calculation was conducted to obtain QSPR equation. Theoretical predictors or independent variable which have an influence on the value of CMC are used to construct QSPR equation. The theoretical predictors are classified into predictor of electronic properties, solubility and steric. A total of 108 experimental CMC belongs to sulfonate-based surfactant are calculated their theoretical predictors and analyzed by multiple linear regression. The QSPR equation which is obtained from this study contains the important theoretical predictors. They are solubility properties, molecular weight, molecular size and net charge of carbon atom in the polar part of sulfonate-based surfactant. This QSPR equation could be used to predict the CMC value of the novel sulfonate-based surfactant.

Keywords: surfactant, sulfonate, QSPR, ZINDO/1**INTRODUCTION**

Surfactant or surface active agent is an important substance. Surfactant has hydrophilic and hydrophobic part. So, it is active on the surface of two immiscible materials and can make the two immiscible materials become one phase (Haftka, 2016). Surfactant also can decrease the surface tension. Surfactant is divided into 4 types such as cationic, anionic, non-ionic and amphoteric surfactant. The surfactant that has a lot of application is anionic surfactant. It can be used as food additive, emulsifier agent at drug industry and also detergent (Paulo, 2017). Since the anionic surfactant has a lot of applications, it needs the development to produce qualified anionic surfactant. A good surfactant has low value of Critical Micelle Concentration (CMC). The CMC is the value of surfactant concentration where the surfactant molecules start to turn into a micelle. A micelle is an aggregation structure of some surfactant molecules in liquid system.

The CMC value becomes important data. The CMC value determination is not easy experimentally (Zhu, 2014). It could be conducted by plotting a graph of a suitable physical property as a function of surfactant concentration. Plotting a graph of a suitable physical property as a function of

surfactant concentration. An abrupt change of slope marks the CMC. The determination of CMC value is difficult, since the change in slope occurs over a more or less narrow range of concentrations and also need a lot of surfactant quantity.

Computational chemistry is proven as useful technique to predict the activity (Iswanto, 2011), synthesis the novel molecule such as anticancer molecule and even the melting point of carbocyclic nitro aromatic compounds (Elidrissi, 2017). This computational chemistry approach is quantitative structure-property relationship (QSPR) analysis (Huibers, 1999). This work is conducted in order to obtain the QSPR equation based on semi empirical ZINDO/1 calculation. This method is the most appropriate method for energy and determining structures of molecules with first or second row metal transition. The experimental data which are used to obtain the QSPR equation are CMC value of sulfonate-based surfactant. The QSPR equation or mathematical equation could be applied to predict the CMC value of novel sulfonate-based surfactant.

EXPERIMENTAL SECTION**Materials**

The experimental Critical Micelle Concentration (CMC) data and molecular formula of 108 anionic

sulfonate-based surfactants (Table 1) (Huibers, 1999)

Instrumentation

The computational chemistry software for calculating the geometry optimization and theoretical

predictor or independent variable of sulfonate-based Surfactant was HyperChem 8.0 (Hypercube, 2007). Multiple Linear Regression (MLR) used as statistical analysis to obtain the Quantitative Structure-Property Relationship (QSPR) equation.

Table 1. Sulfonate surfactants and their experimental CMC value (Log CMC form) (Huibers, 1999)

Compound	Log CMC	Compound	Log CMC	Compound	Log CMC
C ₁₀ PhSO ₃ Na	-2.509	C ₇ C(C ₇)SO ₃ Na	-2.144	C ₁₅ C(OH)C ₂ SO ₃ Na	-3.420
C ₁₂ PhSO ₃ Na	-2.921	C ₇ PhSO ₃ Na	-1.582	C ₁₀ EOSO ₃ Na	-1.787
C ₈ C(C ₂)PhSO ₃ Na	-2.500	C ₈ PhSO ₃ Na	-1.907	C ₁₀ C(C(OH))SO ₃ Na	-1.787
C ₇ C(C ₃)PhSO ₃ Na	-2.400	C ₆ C(C ₂)PhSO ₃ Na	-1.967	C ₁₂ C(C(OH))SO ₃ Na	-2.432
C ₆ C(C ₄)PhSO ₃ Na	-2.300	C ₈ C(C)PhSO ₃ Na	-2.303	C ₆ CO ₂ CSO ₃ Na	-0.733
C ₅ C(C ₅)PhSO ₃ Na	-2.250	C ₇ C(C ₂)PhSO ₃ Na	-2.200	C ₈ CO ₂ CSO ₃ Na	-1.144
C ₇ C(C ₄)PhSO ₃ Na	-2.570	C ₅ C(C ₄)PhSO ₃ Na	-2.047	C ₁₀ CO ₂ CSO ₃ Na	-1.621
C ₁₀ C(C ₂)PhSO ₃ Na	-3.000	C ₉ C(C)PhSO ₃ Na	-2.721	C ₁₄ CO ₂ C ₂ SO ₃ Na	-3.046
C ₉ C(C ₃)PhSO ₃ Na	-2.900	C ₁₀ C(C)PhSO ₃ Na	-2.692	C ₁₀ C(CO ₂ C)SO ₃ Na	-1.986
C ₈ C(C ₄)PhSO ₃ Na	-2.780	C ₉ C(C ₂)PhSO ₃ Na	-2.606	C ₁₂ C(CO ₂ C)SO ₃ Na	-2.583
C ₇ C(C ₅)PhSO ₃ Na	-2.700	C ₈ C(C ₃)PhSO ₃ Na	-2.721	C ₁₄ C(CO ₂ C)SO ₃ Na	-3.398
C ₆ C(C ₆)PhSO ₃ Na	-2.600	C ₆ C(C ₅)PhSO ₃ Na	-2.585	C ₁₄ C(CO ₂ C ₂)SO ₃ Na	-3.509
C ₁₂ C(C)PhSO ₃ Na	-3.390	C ₁₁ C(C)PhSO ₃ Na	-3.208	C ₁₄ C(CO ₂ C ₃)SO ₃ Na	-3.964
C ₁₁ C(C ₂)PhSO ₃ Na	-3.280	C ₁₃ C(C)PhSO ₃ Na	-3.577	C ₁₆ C(CO ₂ C)SO ₃ Na	-4.000
C ₁₀ C(C ₃)PhSO ₃ Na	-3.150	C ₉ C(OH)C ₂ SO ₃ Na	-1.606	C ₁₆ C(CO ₂ C ₂)SO ₃ Na	-4.106
C ₉ C(C ₄)PhSO ₃ Na	-3.050	C ₉ C(OC)C ₂ SO ₃ Na	-2.118	C ₁₆ C(CO ₂ C ₃)SO ₃ Na	-4.899
C ₈ C(C ₅)PhSO ₃ Na	-2.900	C ₉ C(OC ₂)C ₂ SO ₃ Na	-2.301	C ₁₆ C(CO ₂ C(C)C)SO ₃ Na	-4.569
C ₇ C(C ₆)PhSO ₃ Na	-2.800	C ₉ C(OC ₃)C ₂ SO ₃ Na	-2.420	C ₄ C(C ₂)CCO ₂ CC(SO ₃ Na)CO ₂	-2.566
C ₈ CO ₂ C ₂ SO ₃ Na	-1.312	C ₉ C(OC(C)C)C ₂ SO ₃ Na	-2.458	C ₄ CO ₂ C(SO ₃ Na)CCO ₂ C ₄	-0.663
C ₁₀ CO ₂ C ₂ SO ₃ Na	-1.883	C ₉ C(OC ₄)C ₂ SO ₃ Na	-2.817	C ₅ CO ₂ C(SO ₃ Na)CCO ₂ C ₅	-1.239
C ₁₂ CO ₂ C ₂ SO ₃ Na	-2.523	C ₉ C(OC ₆)C ₂ SO ₃ Na	-3.185	C ₆ CO ₂ C(SO ₃ Na)CCO ₂ C ₆	-1.817
C ₆ SO ₃ Na	-0.496	C ₉ C(OC ₈)C ₂ SO ₃ Na	-3.922	C ₈ CO ₂ C(SO ₃ Na)CCO ₂ C ₈	-3.131
C ₈ SO ₃ Na	-0.796	C ₉ C(OCC(C ₂)C ₄)C ₂ SO ₃ Na	-3.501	C ₅ PhePhSO ₃ Na	-1.817
C ₁₀ SO ₃ Na	-1.398	C ₉ C(OPh)C ₂ SO ₃ Na	-2.708	C ₈ PhePhSO ₃ Na	-2.817
C ₁₂ SO ₃ Na	-1.959	C ₉ C(O)C ₂ SO ₃ Na	-1.541	C ₁₀ PhePhSO ₃ Na	-3.445
C ₁₃ SO ₃ Na	-2.421	C ₁₁ C(OH)C ₂ SO ₃ Na	-2.199	C ₁₂ hePhSO ₃ Na	-4.164
C ₁₄ SO ₃ Na	-2.602	C ₁₁ C(OC ₂ OH)C ₂ SO ₃ Na	-3.445	C ₁₄ PhePhSO ₃ Na	-4.759
C ₁₅ SO ₃ Na	-3.139	C ₁₁ C(EO ₂)C ₂ SO ₃ Na	-2.922	C ₁₂ AlaPhSO ₃ Na	-3.351
C ₁₆ SO ₃ Na	-3.131	C ₁₁ C(OPh)C ₂ SO ₃ Na	-3.641	C ₁₁ C(morpholino)C ₂ SO ₃ Na	-3.106
C ₁₇ SO ₃ Na	-3.635	C ₁₁ C(OPhC ₁₃)C ₂ SO ₃ Na	-4.787	C ₁₁ C(piperidino)C ₂ SO ₃ Na	-3.310
C ₁₀ C=CSO ₃ Na	-1.886	C ₁₁ C(NC ₂)C ₂ SO ₃ Na	-2.964		
C ₁₂ C=CSO ₃ Na	-2.569	C ₁₁ C(NC ₃)C ₂ SO ₃ Na	-3.200		
C ₁₄ C=CSO ₃ Na	-3.215	C ₁₁ C(NC ₄)C ₂ SO ₃ Na	-3.708		
C ₁₆ C=CSO ₃ Na	-3.745	C ₁₁ C(O)C ₂ SO ₃ Na	-2.174		
C ₁₀ C(C)SO ₃ Na	-1.827	C ₁₃ C(OH)C ₂ SO ₃ Na	-2.839		
C ₉ C(C ₂)SO ₃ Na	-1.730	C ₁₃ C(OC)C ₂ SO ₃ Na	-3.472		
C ₈ C(C ₃)SO ₃ Na	-1.635	C ₁₃ C(OC ₃)C ₂ SO ₃ Na	-4.089		
C ₇ C(C ₄)SO ₃ Na	-1.548	C ₁₃ C(OC ₄)C ₂ SO ₃ Na	-4.458		
C ₆ C(C ₅)SO ₃ Na	-1.442	C ₁₃ C(O)C ₂ SO ₃ Na	-2.735		

Statistical Package for the Social Sciences (SPSS) ver 25.0 was used. All calculations and analysis was performed by 1 Personal Computer (PC). The specifications of this PC are intel Core2Quad processor, 3 MB of RAM, 500 GB of harddisk and windows 7 operating system, and also MS office.

Procedure

Molecular Modeling of Surfactant

Each compound is drawn into a three dimensional (3D) model of compound in the Hyperchem Release 8.0 molecular modeling program. Then, each model is saved into HIN file format (Hypercube, 2007).

Generating Theoretical Predictors

The predictors in this study are theoretical predictors that are generated by computational chemistry calculation. They consist of the physicochemical properties of sulfonate-based anionic surfactant. The predictors are used as independent variables in statistical analysis (Iswanto, 2011). They are assumed as an important factors on micelle forming and physical properties of surfactant. They are the charge of Carbon atom at polar and non-polar side of hydrocarbon chain of surfactant compound, dipole moment, partition coefficient (Log P), polarizability, refractive index, molecular weight, volume and surface area of van der Waals or QSAR properties calculations. The polar side Carbon atom (C_{polar}) is the Carbon atom which binds the sulfonate functional group. The non polar side Carbon atom ($C_{\text{non polar}}$) is the farthest Carbon atom from the C_{polar} atom in the main hydrocarbon chain of sulfonate-based surfactant.

The charge of Carbon atom and dipole moment are obtained by Geometry Optimization calculation of surfactant structure. The other predictors or independent variables are obtained by QSAR properties calculations. All of the predictor calculations are calculated by Hyperchem Release 8.0. The calculation used semiempirical ZINDO/1 method, RHF spin pairing, the lowest state, Polak-Ribiere optimization algorithm and $0.001 \text{ kcal.}\text{\AA}^{-1}.\text{mole}^{-1}$ of RMS Gradient (Iswanto, 2011). Dependent variable or symbol with P in the Equation (1.1) in this study is the log of experimental Critical Micelle Concentration (Log CMC) of the sulfonate-based anionic surfactant.

Constructing QSPR Equation

Construction of QSPR equation is conducted by Multiple Linear Regression (MLR) statistical analysis based on backward method. All of the theoretical predictors are used as independent variables and the log CMC data are used as dependent variables. By looking to the quality of statistical parameters, the selected QSPR equation is chosen. The general form of QSPR equation is described by **Equation 1**.

$$P = C + b_1.x_1 + b_2.x_2 + b_3.x_3 + \dots \quad (1)$$

Where: P = compound's property

C = a constant

x_n = n^{th} variable = n^{th} predictor

b_n = a coefficient of n^{th} predictor

RESULTS AND DISCUSSION

Molecular Modeling

Molecular modeling is important step for describing the chemical system in computational chemistry approach. The detailed checking of three dimensional structure of the sulfonate surfactant model is also important. If the structure is not correctly represented, the calculation of predictors would also give wrong results. Modeling chemical system is one of important things besides choosing calculation method in computational chemistry study (Iswanto, 2011). The result from modeling step is a three dimensional coordinate of the compound. This three dimensional coordinate is applied for calculation of theoretical predictor.

Theoretical Predictors

Geometry optimization of sulfonate-based surfactant structure was conducted to obtaining the theoretical predictors. The geometry optimization structure calculation used semiempirical ZINDO/1 method. The semiempirical ZINDO/1 method is the electronic structure calculation method that neglects the electron interaction integral. The calculation is not time consuming and electronic structure data can be easily obtained. The ZINDO/1 method is suitable for bulky structure of organic compound which consists of metal ion such as sulfonate-based surfactant. The semiempirical ZINDO/1 calculation method is already applied in corrosion inhibition of steel with some amino acids (Sikachina, 2018).

The theoretical predictors are the physicochemical properties which influence the value of Critical Micelle Concentration (Delsy, 2017). Results of theoretical predictors calculation show the theoretical characteristics of sulfonate-based anionic surfactant. The Carbon (C) atom net charges at polar and non-polar side of main hydrocarbon chain and also dipole moment of surfactant describe the electronic properties, where as partition coefficient and molecular weight reflect the solubility and sterical properties of surfactant compounds, respectively. In the atomic net charges value, the various value is observed at the C_{polar} and $C_{\text{non polar}}$. These significant differences in these C atom net charges are due to the difference in functional group or substituent which is binded at them.

If sulfonate functional group is binded by phenyl and also branched alkene, the C_{polar} will have positive net charge. In contrast, if sulfonate functional group is binded by linear alkene, the C_{polar} will have negative net charge. For instance, C_{polar} of $C_8C(C_3)PhSO_3Na$ and C_8SO_3Na are 0.0778 and -0.060 coulomb, respectively. The C_{polar} of

$C_8C(C_3)PhSO_3Na$ has positive charge due to the attractive electron functional group of sulfonate. In the case of C_8SO_3Na , the linear hydrocarbon could push the electron to the C_{polar} and the C_{polar} charge becomes negative (Iswanto, 2011). In the one sulfonate-based surfactant molecule the charge of C_{polar} is more positive than $C_{non\ polar}$. This is occurred due to the attractive electron functional group of sulfonate at the C_{polar} .

The dipole moment values are also different among the sulfonate-based surfactant. The highest value is observed in $C_8PhePhSO_3Na$ compound. Its dipole moment value is 27.477 D. This surfactant has various electronegative atom in the molecular structure and large of size. The symbol Phe and Ph represent the Phenylalanine amino acid and Phenyl, respectively. The lowest dipole moment value is observed in $C_7C(C_6)PhSO_3Na$ compound. The value is 3.381 D. This compound does not have various electronegative atom. Dipole moment value is influenced by the difference in electronegativity among the atoms and the geometry of certain functional group in the structure (Iswanto, 2011). Solubility property is described by the Log P value. The highest dan the lowest Log P value are found in $C_{11}C(OPhC_{13})C_2SO_3Na$ and $C_6CO_2CSO_3Na$, respectively.

Based on molecular structure, the $C_6CO_2CSO_3Na$ compound is more polar and more soluble in water. Solubility property is influenced by the property of functional group in the compound. The sterical properties are observed by refractive index, molecular weight, volume and surface area of van der Waals. The surfactant which has various atom and large size of molecule, will have high value of sterical properties.

The QSPR Equation Candidates

The theoretical predictors which calculated by semiempirical ZINDO/1 calculation are applied for obtaining the QSPR equation. Statistical analysis is applied for obtaining this equation. Statistical analysis which is applied for obtaining QSPR equation is regression analysis. The dependent variables are 108 CMC experimental data. The independent variables are the theoretical predictors. The number of dependent and independent variables are enough to conducting the multiple linear regression (MLR) analysis. The MLR is the simple regression analysis for multiple independent variable. In one step of statistical calculation, the QSPR equation will be obtained (Delsy, 2017).

Before MLR analysis conducted, the dependent variables are divided into 2 categories. The first category is fitting data which are used to obtaining the QSPR equation. The second category is the testing data which are used to selecting the best QSPR equation. Based on the random selection or used RAND function of Microsoft Excel calculation, the fitting data are 70 sulfonate surfactant compounds and the testing data are 38 compounds (Table 2).

Table 3 shows the important theoretical predictors in the six QSPR equation candidates. These candidates are obtained by Backward method of MLR statistical analysis. The six QSPR equation candidates have almost similar correlation coefficient (r) value. Further step is needed for selecting the best QSPR equation. Selecting the best QSPR equation form the six candidates is conducted by the 38 sulfonate surfactants as testing data.

Table 2. Sulfonate surfactants that used to be testing data or testing compounds for QSPR equation selection

Testing compounds or testing data for QSPR equation		
$C_{12}C(C)PhSO_3Na$	$C_9C(OC_2)C_2SO_3Na$	$C_8C(C_5)PhSO_3Na$
$C_{11}C(OC_2OH)C_2SO_3Na$	$C_4C(C_2)CCO_2CC(SO_3Na)CO_2CC(C_2)C_4$	$C_9C(C_3)PhSO_3Na$
$C_{10}C=CSO_3Na$	$C_{11}C(C)PhSO_3Na$	$C_{14}C(CO_2C)SO_3Na$
$C_{12}C(CO_2C)SO_3Na$	$C_9C(C_2)PhSO_3Na$	$C_{14}PhePhSO_3Na$
$C_{10}C(C)SO_3Na$	$C_{15}C(OH)C_2SO_3Na$	$C_{13}C(C)PhSO_3Na$
$C_9C(OH)C_2SO_3Na$	$C_{13}C(O)C_2SO_3Na$	$C_7C(C_6)PhSO_3Na$
$C_{12}SO_3Na$	$C_{14}C(CO_2C_2)SO_3Na$	$C_{11}C(EO_2)C_2SO_3Na$
$C_6C(C_4)PhSO_3Na$	$C_9C(OC)C_2SO_3Na$	$C_{13}C(OC)C_2SO_3Na$
$C_{13}C(OH)C_2SO_3Na$	$C_{12}CO_2C_2SO_3Na$	$C_{12}C(C(OH))SO_3Na$
$C_7C(C_2)PhSO_3Na$	$C_{11}C(NC_3)C_2SO_3Na$	$C_9C(C_2)SO_3Na$
$C_{11}C(Morpholino)C_2SO_3Na$	$C_{11}C(Piperidino)C_2SO_3Na$	$C_{11}C(OPh)C_2SO_3Na$
$C_{11}C(OPhC_{13})C_2SO_3Na$	$C_4CO_2(SO_3Na)CCO_2C_4$	$C_{16}C(CO_2C_3)SO_3Na$
$C_9C(OPh)C_2SO_3Na$	$C_9C(C_4)PhSO_3Na$	

Table 3. The important theoretical predictors or independent variable in the QSPR equation candidates based on statistical calculation of multiple linear regression Backward method. These candidates are obtained by fitting the 70 CMC experimental data of sulfonate-based surfactant compounds.

Model or Equation	Independent Variable	r	r ²	SE	F _{cal} /F _{table}
1	qC ₁ , qC ₂ , R _d , Log P, Polarizability, dipole moment, MW, A _{v_{vdw}} , V _{v_{vdw}}	0.965	0.931	0.259898	44.193954
2	qC ₁ , qC ₂ , Log P, Polarizability, dipole moment, MW, A _{v_{vdw}} , V _{v_{vdw}}	0.965	0.931	0.257764	25.771324
3	qC ₁ , qC ₂ , Log P, Polarizability, dipole moment, MW, V _{v_{vdw}}	0.965	0.931	0.256297	29.800466
4	qC ₁ , qC ₂ , Log P, V _{v_{vdw}} , MW, dipole moment	0.964	0.930	0.255924	34.857340
5	qC ₁ , qC ₂ , Log P, V _{v_{vdw}} , MW	0.964	0.929	0.255685	41.889234
6	qC ₁ , Log P, V _{v_{vdw}} , MW	0.963	0.927	0.256314	52.054614

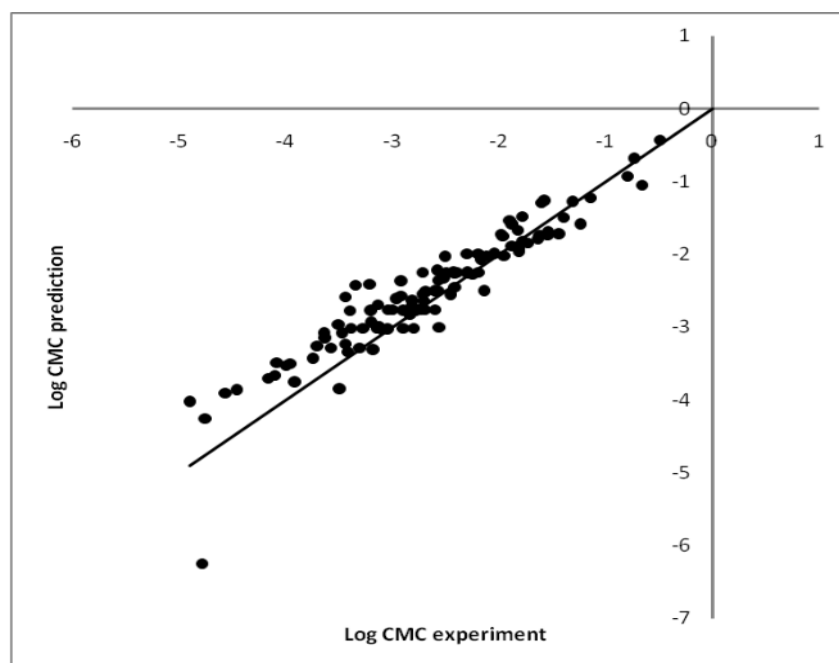


Figure 1. Plot the Log CMC prediction of sulfonate-based surfactants which are calculated by QSPR equation based on semiempirical ZINDO/1 against their Log CMC experiment (dot). The straight line represents if all of the Log CMC prediction values equal to Log CMC experiment.

The Final QSPR equation

Selection of the best QSPR equation is conducted with evaluation of the statistical parameters such as r (correlation coefficient), r^2 (determination coefficient), F (variance analysis), SE (standard error) and $PRESS$ (Prediction Residual Sum of Squares) (Iswanto, 2011). Based on the evaluation process against the six QSPR equation candidates, the 6th QSPR candidate is the best equation. It has good quality of statistical parameters (Table 2). It also has the smallest of $PRESS$ value. It means that the 6th QSPR candidate has the best quality for predicting the theoretical CMC value of sulfonate-based surfactant.

The final QSPR equation is obtained by multiple linear regression enter method to the fitting and

testing data or to the 108 CMC of sulfonate-based surfactant. The final QSPR equation consists of 4 important theoretical predictors. There are the charge of C_{polar} (qC_1), partition coefficient of *n*-octanol/water (Log P), van der Waals volume (V_{vdw}) and molecular weight (MW). The final QSPR equation is:

$$\text{Log CMC} = 1.519 - 0.155\text{LogP} + 0.010\text{MW} - 0.021V_{vdw} + 2.336qC_1$$

This final QSPR equation has some characteristics. They are the number of data (n) = 108, r = 0.937, r^2 = 0.878, SE = 0.3236, and F_{calc}/F_{table} = 75.862 and $PRESS$ = 10.9394.

The final QSPR equation has good correlation (r value) between Log CMC value and the 4 theoretical predictors or independent variable. This QSPR

equation could explain over 87.8% variation of data (see r^2 value). Small error is also showed in this final QSPR equation (see SE value). This final QSPR equation is also accepted at 95% confidence level because its F_{calc}/F_{table} is more than 1.00 (Iswanto, 2011). Based on the final QSPR equation, the most important theoretical predictor is the charge of C_{polar} . This theoretical predictor has the highest value of coefficient. The capability of the QSPR equation in order to predict the CMC value is good. Most of the CMC prediction value is close to the CMC experiment (Figure 1).

CONCLUSIONS

Computational chemistry calculation based on semiempirical ZINDO/1 method is good compromising between the accuracy and the capability of computer. The QSPR equation could be obtained by conducting the QSPR analysis against the CMC experimental value of sulfonate-based surfactants. The most important theoretical predictor could be shown in the highest coefficient value of its in the QSPR equation.

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Quantitative Structure-Property Relationship Analysis Against Critical Micelle Concentration of Sulfonate-Based Surfactant Based on Semiempirical Zindo/1 Calculation

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ABSTRACT. Development of anionic surfactant compound is very important because the anionic surfactant class is widely used in people's lives. For instance, anionic surfactants are used as food additives and detergents. The novel compound of sulfonate-based surfactant or proposed compound has predicted the Critical Micelle Concentration (CMC) value of experiment. Quantitative Structure-Property Relationship (QSPR) analysis based on semiempirical ZINDO/1 calculation was conducted to obtain QSPR equation. Theoretical predictors or independent variable which have an influence on the value of CMC are used to construct QSPR equation. The theoretical predictors are classified into predictor of electronic properties, solubility and steric. A total of 108 experimental CMC belongs to sulfonate-based surfactant are calculated their theoretical predictors and analyzed by multiple linear regression. The QSPR equation which is obtained from this study contains the important theoretical predictors. They are solubility properties, molecular weight, molecular size and net charge of carbon atom in the polar part of sulfonate-based surfactant. This QSPR equation could be used to predict the CMC value of the novel sulfonate-based surfactant.

Keywords: surfactant, sulfonate, QSPR, ZINDO/1

INTRODUCTION

Surfactant or surface active agent is an important substance. Surfactant has hydrophilic and hydrophobic part. So, it is active on the surface of two immiscible materials and can make the two immiscible materials become one phase (Haftka, 2016). Surfactant also can decrease the surface tension. Surfactant is divided into 4 types such as cationic, anionic, non-ionic and amphoteric surfactant. The surfactant that has a lot of application is anionic surfactant. It can be used as food additive, emulsifier agent at drug industry and also detergent (Paulo, 2017). Since the anionic surfactant has a lot of applications, it needs the development to produce a modified anionic surfactant. A good surfactant has low value of Critical Micelle Concentration (CMC). The CMC is the value of surfactant concentration where the surfactant molecules start to turn into a micelle. A micelle is an aggregation structure of some surfactant molecules in liquid system.

The CMC value becomes important data. The CMC value determination is not easy experimentally (Zhu, 2014). It could be conducted by plotting a graph of a suitable physical property as a function of surfactant concentration. Plotting a graph of a suitable physical property as a function of

surfactant concentration. An abrupt change of slope marks the CMC. The determination of CMC value is difficult, since the change in slope occurs over a more or less narrow range of concentrations and also need a lot of surfactant quantity.

Computational chemistry is proven as useful technique to predict the activity (Iswanto, 2011), synthesis the novel molecule such as anticancer molecule and even the melting point of carbonic nitro aromatic compounds (Elidrissi, 2017). This computational chemistry approach is quantitative structure-property relationship (QSPR) analysis (Huibers, 1999). This work is conducted in order to obtain the QSPR equation based on semi empirical ZINDO/1 calculation. This method is the most appropriate method for energy and determining structures of molecules with first or second row metal transition. The experimental data which are used to obtain the QSPR equation are CMC value of sulfonate-based surfactant. The QSPR equation or mathematical equation could be applied to predict the CMC value of novel sulfonate-based surfactant.

EXPERIMENTAL SECTION

Materials

The experimental Critical Micelle Concentration (CMC) data and molecular formula of 108 anionic

sulfonate-based surfactants (Table 1) (Huibers, 1999)

Instrumentation

The computational chemistry software for calculating the geometry optimization and theoretical

predictor or independent variable of sulfonate-based Surfactant was HypercheRelease 8.0 (Hypercube, 2007). Multiple Linear Regression (MLR) used as statistical analysis to obtain the Quantitative Structure-Property Relationship (QSPR) equation.

Table 1. Sulfonate surfactants and their experimental CMC value (Log CMC form) (Huibers, 1999)

Compound	Log CMC	Compound	Log CMC	Compound	Log CMC
C ₁₀ PhSO ₃ Na	-2.509	C ₇ C(C ₇)SO ₃ Na	-2.144	C ₁₅ C(OH)C ₂ SO ₃ Na	-3.420
C ₁₂ PhSO ₃ Na	-2.921	C ₇ PhSO ₃ Na	-1.582	C ₁₀ EOSO ₃ Na	-1.787
C ₈ C(C ₂)PhSO ₃ Na	-2.500	C ₈ PhSO ₃ Na	-1.907	C ₁₀ C(C(OH))SO ₃ Na	-1.787
C ₇ C(C ₃)PhSO ₃ Na	-2.400	C ₆ C(C ₂)PhSO ₃ Na	-1.967	C ₁₂ C(C(OH))SO ₃ Na	-2.432
C ₆ C(C ₄)PhSO ₃ Na	-2.300	C ₈ C(C)PhSO ₃ Na	-2.303	C ₆ CO ₂ CSO ₃ Na	-0.733
C ₅ C(C ₅)PhSO ₃ Na	-2.250	C ₇ C(C ₂)PhSO ₃ Na	-2.200	C ₈ CO ₂ CSO ₃ Na	-1.144
C ₇ C(C ₄)PhSO ₃ Na	-2.570	C ₅ C(C ₄)PhSO ₃ Na	-2.047	C ₁₀ CO ₂ CSO ₃ Na	-1.621
C ₁₀ C(C ₂)PhSO ₃ Na	-3.000	C ₉ C(C)PhSO ₃ Na	-2.721	C ₁₄ CO ₂ C ₂ SO ₃ Na	-3.046
C ₉ C(C ₃)PhSO ₃ Na	-2.900	C ₁₀ C(C)PhSO ₃ Na	-2.692	C ₁₀ C(CO ₂ C)SO ₃ Na	-1.986
C ₈ C(C ₄)PhSO ₃ Na	-2.780	C ₉ C(C ₂)PhSO ₃ Na	-2.606	C ₁₂ C(CO ₂ C)SO ₃ Na	-2.583
C ₇ C(C ₅)PhSO ₃ Na	-2.700	C ₈ C(C ₃)PhSO ₃ Na	-2.721	C ₁₄ C(CO ₂ C)SO ₃ Na	-3.398
C ₆ C(C ₆)PhSO ₃ Na	-2.600	C ₆ C(C ₅)PhSO ₃ Na	-2.585	C ₁₄ C(CO ₂ C ₂)SO ₃ Na	-3.509
C ₁₂ C(C)PhSO ₃ Na	-3.390	C ₁₁ C(C)PhSO ₃ Na	-3.208	C ₁₄ C(CO ₂ C ₃)SO ₃ Na	-3.964
C ₁₁ C(C ₂)PhSO ₃ Na	-3.280	C ₁₃ C(C)PhSO ₃ Na	-3.577	C ₁₆ C(CO ₂ C)SO ₃ Na	-4.000
C ₁₀ C(C ₃)PhSO ₃ Na	-3.150	C ₉ C(OH)C ₂ SO ₃ Na	-1.606	C ₁₆ C(CO ₂ C ₂)SO ₃ Na	-4.106
C ₉ C(C ₄)PhSO ₃ Na	-3.050	C ₉ C(OC)C ₂ SO ₃ Na	-2.118	C ₁₆ C(CO ₂ C ₃)SO ₃ Na	-4.899
C ₈ C(C ₅)PhSO ₃ Na	-2.900	C ₉ C(OC ₂)C ₂ SO ₃ Na	-2.301	C ₁₆ C(CO ₂ C(C)C)SO ₃ Na	-4.569
C ₇ C(C ₆)PhSO ₃ Na	-2.800	C ₉ C(OC ₃)C ₂ SO ₃ Na	-2.420	C ₄ C(C ₂)CCO ₂ CC(SO ₃ Na)CO ₂	-2.566
C ₈ CO ₂ C ₂ SO ₃ Na	-1.312	C ₉ C(OC(C)C)C ₂ SO ₃ Na	-2.458	C ₄ CO ₂ C(SO ₃ Na)CCO ₂ C ₄	-0.663
C ₁₀ CO ₂ C ₂ SO ₃ Na	-1.883	C ₉ C(OC ₄)C ₂ SO ₃ Na	-2.817	C ₅ CO ₂ C(SO ₃ Na)CCO ₂ C ₅	-1.239
C ₁₂ CO ₂ C ₂ SO ₃ Na	-2.523	C ₉ C(OC ₆)C ₂ SO ₃ Na	-3.185	C ₆ CO ₂ C(SO ₃ Na)CCO ₂ C ₆	-1.817
C ₆ SO ₃ Na	-0.496	C ₉ C(OC ₈)C ₂ SO ₃ Na	-3.922	C ₈ CO ₂ C(SO ₃ Na)CCO ₂ C ₈	-3.131
C ₈ SO ₃ Na	-0.796	C ₉ C(OCC(C ₂)C ₄)C ₂ SO ₃ Na	-3.501	C ₅ PhePhSO ₃ Na	-1.817
C ₁₀ SO ₃ Na	-1.398	C ₉ C(OPh)C ₂ SO ₃ Na	-2.708	C ₈ PhePhSO ₃ Na	-2.817
C ₁₂ SO ₃ Na	-1.959	C ₉ C(O)C ₂ SO ₃ Na	-1.541	C ₁₀ PhePhSO ₃ Na	-3.445
C ₁₃ SO ₃ Na	-2.421	C ₁₁ C(OH)C ₂ SO ₃ Na	-2.199	C ₁₂ hePhSO ₃ Na	-4.164
C ₁₄ SO ₃ Na	-2.602	C ₁₁ C(OC ₂ OH)C ₂ SO ₃ Na	-3.445	C ₁₄ PhePhSO ₃ Na	-4.759
C ₁₅ SO ₃ Na	-3.139	C ₁₁ C(EO ₂)C ₂ SO ₃ Na	-2.922	C ₁₂ AlaPhSO ₃ Na	-3.351
C ₁₆ SO ₃ Na	-3.131	C ₁₁ C(OPh)C ₂ SO ₃ Na	-3.641	C ₁₁ C(morpholino)C ₂ SO ₃ Na	-3.106
C ₁₇ SO ₃ Na	-3.635	C ₁₁ C(OPhC ₁₃)C ₂ SO ₃ Na	-4.787	C ₁₁ C(piperidino)C ₂ SO ₃ Na	-3.310
C ₁₀ C=CSO ₃ Na	-1.886	C ₁₁ C(NC ₂)C ₂ SO ₃ Na	-2.964		
C ₁₂ C=CSO ₃ Na	-2.569	C ₁₁ C(NC ₃)C ₂ SO ₃ Na	-3.200		
C ₁₄ C=CSO ₃ Na	-3.215	C ₁₁ C(NC ₄)C ₂ SO ₃ Na	-3.708		
C ₁₆ C=CSO ₃ Na	-3.745	C ₁₁ C(O)C ₂ SO ₃ Na	-2.174		
C ₁₀ C(C)SO ₃ Na	-1.827	C ₁₃ C(OH)C ₂ SO ₃ Na	-2.839		
C ₉ C(C ₂)SO ₃ Na	-1.730	C ₁₃ C(OC)C ₂ SO ₃ Na	-3.472		
C ₈ C(C ₃)SO ₃ Na	-1.635	C ₁₃ C(OC ₃)C ₂ SO ₃ Na	-4.089		
C ₇ C(C ₄)SO ₃ Na	-1.548	C ₁₃ C(OC ₄)C ₂ SO ₃ Na	-4.458		
C ₆ C(C ₅)SO ₃ Na	-1.442	C ₁₃ C(O)C ₂ SO ₃ Na	-2.735		

Statistical Package for the Social Sciences (SPSS) ver 25.0 was used. All calculations and analysis was performed by 1 Personal Computer (PC). The specifications of this PC are intel Core2Quad processor, 3 MB of RAM, 500 GB of harddisk and windows 7 operating system, and also MS office.

Procedure

Molecular Modeling of Surfactant

Each compound is drawn into a three dimensional (3D) model of compound in the Hyperchem Release 8.0 molecular modeling program. Then, each model is saved into HIN file format (Hypercube, 2007).

Generating Theoretical Predictors

The predictors in this study are theoretical predictors that are generated by computational chemistry calculation. They consist of the physicochemical properties of sulfonate-based anionic surfactant. The predictors are used as independent variables in statistical analysis (Iswanto, 2011). They are assumed as an important factors on micelle forming and physical properties of surfactant. They are the charge of Carbon atom at polar and non-polar side of hydrocarbon chain of surfactant compound, dipole moment, partition coefficient (Log P), polarizability, refractive index, molecular weight, volume and surface area of van der Waals or QSAR properties calculations. The polar side Carbon atom (C_{polar}) is the Carbon atom which binds the sulfonate functional group. The non polar side Carbon atom ($C_{non\ polar}$) is the farthest Carbon atom from the C_{polar} atom in the main hydrocarbon chain of sulfonate-based surfactant.

The charge of Carbon atom and dipole moment are obtained by Geometry Optimization calculation of surfactant structure. The other predictors or independent variables are obtained by QSAR properties calculations. All of the predictor calculations are calculated by Hyperchem Release 8.0. The calculation used semiempirical ZINDO/1 method, RHF spin pairing, the lowest state, Polak-Ribiere optimization algorithm and 0.001 kcal.Å⁻¹.mole⁻¹ of RMS Gradient (Iswanto, 2011). Dependent variable or symbol with P in the Equation (1.1) in this study is the log of experimental Critical Micelle Concentration (Log CMC) of the sulfonate-based anionic surfactant.

Constructing QSPR Equation

Construction of QSPR equation is conducted by Multiple Linear Regression (MLR) statistical analysis based on backward method. All of the theoretical predictors are used as independent variables and the log CMC data are used as dependent variables. By looking to the quality of statistical parameters, the selected QSPR equation is chosen. The general form of QSPR equation is described by Equation 1.

$$P = C + b_1.x_1 + b_2.x_2 + b_3.x_3 + \dots \quad (1)$$

Where: P = compound's property

C = a constant

x_n = nth variable = nth predictor

b_n = a coefficient of nth predictor

RESULTS AND DISCUSSION

Molecular Modeling

Molecular modeling is important step for describing the chemical system in computational chemistry approach. The detailed checking of three dimensional structure of the sulfonate surfactant model is also important. If the structure is not correctly represented, the calculation of predictors would also give wrong results. Modeling chemical system is one of important things besides choosing calculation method in computational chemistry study (Iswanto, 2011). The result from modeling step is a three dimensional coordinate of the compound. This three dimensional coordinate is applied for calculation of theoretical predictor.

Theoretical Predictors

Geometry optimization of sulfonate-based surfactant structure was conducted to obtaining the theoretical predictors. The geometry optimization structure calculation used semiempirical ZINDO/1 method. The semiempirical ZINDO/1 method is the electronic structure calculation method that neglects the electron interaction integral. The calculation is not time consuming and electronic structure data can be easily obtained. The ZINDO/1 method is suitable for bulky structure of organic compound which consists of metal ion such as sulfonate-based surfactant. The semiempirical ZINDO/1 calculation method is already applied in corrosion inhibition of steel with some amino acids (Sikachina, 2018).

The theoretical predictors are the physicochemical properties which influence the value of Critical Micelle Concentration (Delsy, 2017). Results of theoretical predictors calculation show the theoretical characteristics of sulfonate-based anionic surfactant. The Carbon (C) atom net charges at polar and non-polar side of main hydrocarbon chain and also dipole moment of surfactant describe the electronic properties, where as partition coefficient and molecular weight reflect the solubility and sterical properties of surfactant compounds, respectively. In the atomic net charges value, the various value is observed at the C_{polar} and $C_{non\ polar}$. These significant differences in these C atom net charges are due to the difference in functional group or substituent which is binded at them.

If sulfonate functional group is binded by phenyl and also branched alkene, the C_{polar} will have positive net charge. In contrast, if sulfonate functional group is binded by linear alkene, the C_{polar} will have negative net charge. For instance, $C_8C(C_3)PhSO_3Na$ and C_8SO_3Na are 0.0778 and -0.060 coulomb, respectively. The C_{polar} of

$C_8C(C_3)PhSO_3Na$ has positive charge due to the attractive electron functional group of sulfonate. In the case of C_8SO_3Na , the linear hydrocarbon could push the electron to the C_{polar} and the C_{polar} charge becomes negative (Iswanto, 2011). In the one sulfonate-based surfactant molecule the charge of C_{polar} is more positive than $C_{non\ polar}$. This is occurred due to the attractive electron functional group of sulfonate at the C_{polar} .

The dipole moment values are also different among the sulfonate-based surfactant. The highest value is observed in $C_8PhePhSO_3Na$ compound. Its dipole moment value is 27.477 D. This surfactant has various electronegative atom in the molecular structure and large of size. The symbol Phe and Ph represent the Phenylalanine amino acid and Phenyl, respectively. The lowest dipole moment value is observed in $C_7C(C_6)PhSO_3Na$ compound. The value is 3.381 D. This compound does not have various electronegative atom. Dipole moment value is influenced by the difference in electronegativity among the atoms and the geometry of certain functional group in the structure (Iswanto, 2011). Solubility property is described by the Log P value. The highest dan the lowest Log P value are found in $C_{11}C(OPhC_{13})C_2SO_3Na$ and $C_6CO_2CSO_3Na$, respectively.

Based on molecular structure, the $C_6CO_2CSO_3Na$ compound is more polar and more soluble in water. Solubility property is influenced by the property of functional group in the compound. The sterical properties are observed by refractive index, molecular weight, volume and surface area of van der Waals. The surfactant which has various atom and large size of molecule, will have high value of sterical properties.

The QSPR Equation Candidates

The theoretical predictors which calculated by semiempirical ZINDO/1 calculation are applied for obtaining the QSPR equation. Statistical analysis is applied for obtaining this equation. Statistical analysis which is applied for obtaining QSPR equation is regression analysis. The dependent variables are 108 CMC experimental data. The independent variables are the theoretical predictors. The number of dependent and independent variables are enough to conducting the multiple linear regression (MLR) analysis. The MLR is the simple regression analysis for multiple independent variable. In one step of statistical calculation, the QSPR equation will be obtained (Delsy, 2017).

Before MLR analysis conducted, the dependent variables are divided into 2 categories. The first category is fitting data which are used to obtaining the QSPR equation. The second category is the testing data which are used to selecting the best QSPR equation. Based on the random selection or used RAND function of Microsoft Excel calculation, the fitting data are 70 sulfonate surfactant compounds and the testing data are 38 compounds (Table 2).

Table 3 shows the important theoretical predictors in the six QSPR equation candidates. These candidates are obtained by Backward method of MLR statistical analysis. The six QSPR equation candidates have almost similar correlation coefficient (r) value. Further step is needed for selecting the best QSPR equation. Selecting the best QSPR equation form the six candidates is conducted by the 38 sulfonate surfactants as testing data.

Table 2. Sulfonate surfactants that used to be testing data or testing compounds for QSPR equation selection

Testing compounds or testing data for QSPR equation		
$C_{12}C(C)PhSO_3Na$	$C_9C(OC_2)C_2SO_3Na$	$C_8C(C_5)PhSO_3Na$
$C_{11}C(OC_2OH)C_2SO_3Na$	$C_4C(C_2)CCO_2CC(SO_3Na)CO_2CC(C_2)C_4$	$C_9C(C_3)PhSO_3Na$
$C_{10}C=CSO_3Na$	$C_{11}C(C)PhSO_3Na$	$C_{14}C(CO_2C)SO_3Na$
$C_{12}C(CO_2C)SO_3Na$	$C_9C(C_2)PhSO_3Na$	$C_{14}PhePhSO_3Na$
$C_{10}C(C)SO_3Na$	$C_{15}C(OH)C_2SO_3Na$	$C_{13}C(C)PhSO_3Na$
$C_9C(OH)C_2SO_3Na$	$C_{13}C(O)C_2SO_3Na$	$C_7C(C_6)PhSO_3Na$
$C_{12}SO_3Na$	$C_{14}C(CO_2C_2)SO_3Na$	$C_{11}C(EO_2)C_2SO_3Na$
$C_6C(C_4)PhSO_3Na$	$C_9C(OC)C_2SO_3Na$	$C_{13}C(OC)C_2SO_3Na$
$C_{13}C(OH)C_2SO_3Na$	$C_{12}CO_2C_2SO_3Na$	$C_{12}C(C(OH))SO_3Na$
$C_7C(C_2)PhSO_3Na$	$C_{11}C(NC_3)C_2SO_3Na$	$C_9C(C_2)SO_3Na$
$C_{11}C(Morpholino)C_2SO_3Na$	$C_{11}C(Piperidino)C_2SO_3Na$	$C_{11}C(OPh)C_2SO_3Na$
$C_{11}C(OPhC_{13})C_2SO_3Na$	$C_4CO_2(SO_3Na)CCO_2C_4$	$C_{16}C(CO_2C_3)SO_3Na$
$C_9C(OPh)C_2SO_3Na$	$C_9C(C_4)PhSO_3Na$	

Table 3. The important theoretical predictors or independent variable in the QSPR equation candidates based on statistical calculation of multiple linear regression Backward method. These candidates are obtained by fitting the 70 CMC experimental data of sulfonate-based surfactant compounds.

Model or Equation	Independent Variable	r	r ²	SE	F _{cal} /F _{table}
1	qC ₁ , qC ₂ , R _d , Log P, Polarizability, dipole moment, MW, A _{v_{vdw}} , V _{vdw}	0.965	0.931	0.259898	44.193954
2	qC ₁ , qC ₂ , Log P, Polarizability, dipole moment, MW, A _{v_{vdw}} , V _{vdw}	0.965	0.931	0.257764	25.771324
3	qC ₁ , qC ₂ , Log P, Polarizability, dipole moment, MW, V _{vdw}	0.965	0.931	0.256297	29.800466
4	qC ₁ , qC ₂ , Log P, V _{vdw} , MW, dipole moment	0.964	0.930	0.255924	34.857340
5	qC ₁ , qC ₂ , Log P, V _{vdw} , MW	0.964	0.929	0.255685	41.889234
6	qC ₁ , Log P, V _{vdw} , MW	0.963	0.927	0.256314	52.054614

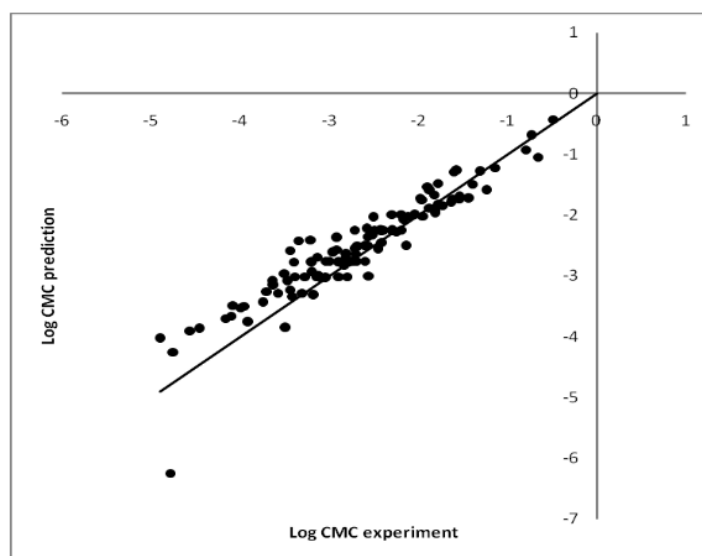


Figure 1. Plot the Log CMC prediction of sulfonate-based surfactants which are calculated by QSPR equation based on semiempirical ZINDO/1 against their Log CMC experiment (dot). The straight line represents if all of the Log CMC prediction values equal to Log CMC experiment.

The Final QSPR equation

Selection of the best QSPR equation is conducted with evaluation of the statistical parameters such as *r* (correlation coefficient), *r*² (determination coefficient), *F* (variance analysis), *SE* (standard error) and *PRESS* (prediction Residual Sum of Squares) (Iswanto, 2011). Based on the evaluation process against the six QSPR equation candidates, the 6th QSPR candidate is the best equation. It has good quality of statistical parameters (Table 2). It also has the smallest of *PRESS* value. It means that the 6th QSPR candidate has the best quality for predicting the theoretical CMC value of sulfonate-based surfactant.

The final QSPR equation is obtained by multiple linear regression enter method to the fitting and

testing data or to the 108 CMC of sulfonate-based surfactant. The final QSPR equation consists of 4 important theoretical predictors. There are the charge of C_{polar} (qC₁), partition coefficient of n-octanol/water (Log P), van der Waals volume (V_{vdw}) and molecular weight (MW). The final QSPR equation is:

$$\text{Log CMC} = 1.519 - 0.155\text{LogP} + 0.010\text{MW} - 0.011\text{V}_{\text{vdw}} + 2.336\text{qC}_1$$

This final QSPR equation has some characteristics. They are the number of data (*n*) = 108, *r* = 0.937, *r*² = 0.878, *SE* = 0.3236, and *F*_{cal}/*F*_{table} = 75.862 and *PRESS* = 10.9394.

The final QSPR equation has good correlation (*r* value) between Log CMC value and the 4 theoretical predictors or independent variable. This QSPR

equation could explain over 87.8% variation of data (see r^2 value). Small error is also showed in this final QSPR equation (see SE value). This final QSPR equation is also accepted at 95% confidence level because its F_{calc}/F_{table} is more than 1.00 (Iswanto, 2011). Based on the final QSPR equation, the most important theoretical predictor is the charge of C_{polar} . This theoretical predictor has the highest value of coefficient. The capability of the QSPR equation in order to predict the CMC value is good. Most of the CMC prediction value is close to the CMC experiment (Figure 1).

CONCLUSIONS

Computational chemistry calculation based on semiempirical ZINDO/1 method is good compromising between the accuracy and the capability of computer. The QSPR equation could be obtained by conducting the QSPR analysis against the CMC experimental value of sulfonate-based surfactants. The most important theoretical predictor could be shown in the highest coefficient value of its in the QSPR equation.

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